

Download Colligative Properties Of Electrolyte Solutions Pdf

In chemistry, colligative properties are properties of solutions that depend on the ratio of the number of solute particles to the number of solvent molecules in a solution, and not on the nature of the chemical species present. The number ratio can be related to the various units for concentration of solutions. Chapter 12 Solutions solute + solvent ! solution speciation? stoichiometry? empirical solubility rules: Which ionic compounds are soluble in water? Freezing-point depression is the decrease of the freezing point of a solvent on addition of a non-volatile solute. Examples include salt in water, alcohol in water, or the mixing of two solids such as impurities into a finely powdered drug. Page 5 of 54 7. Buffer Buffer equations and buffer capacity in general. Buffers in pharmaceutical systems, preparations and stability, buffered isotonic solutions.